

**SAM HOUSTON STATE UNIVERSITY**  
**DEPARTMENT OF CHEMISTRY**  
**CHEMISTRY 1411 – General Chemistry I**  
**Course Syllabus – Fall 2017**

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**MEETING TIME**

Lecture:       Section 08       Chemistry and Forensic Science Building - CFS 121  
                  MWF 4:00 – 4:50 pm

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**INSTRUCTOR INFORMATION**

Dr. Adrian Villalta-Cerdas

e-mail:        axv067@shsu.edu

Office:        CFS 114

Office hours: Tu 2-4 pm, Th 2-5 pm; stop by or other times by appointment.

Office phone: 294 – 2556

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**TEXTBOOK AND MATERIALS**

- 1) Brown, LeMay, Bursten, Murphy, Woodward and Stoltzfus. Chemistry: The Central Science, 13<sup>th</sup> Edition (ISBN-13: 978-0321910417). Older editions are good too, but the 13<sup>th</sup> edition will be our reference book.
- 2) Calculator: recommended type is TI-30X.
- 3) Recommended (if available): portable computer device such as smart phone, iPad or small laptop for viewing PDF and PowerPoint files.
- 4) iClicker Remote or iclicker App (“Reef Polling”) for smartphones.
- 5) Recommended websites:
  - a. <https://openstaxcollege.org/textbooks/chemistry>
  - b. [http://preparatorychemistry.com/Bishop\\_Atoms\\_First.htm](http://preparatorychemistry.com/Bishop_Atoms_First.htm)
  - c. <http://2012books.lardbucket.org/books/principles-of-general-chemistry-v1.0/>

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**COURSE DESCRIPTION**

**Overview**

CHEM1411, General Chemistry I, is the first semester of a two-semester course that offers an introduction to the science of chemistry. In this course we will discuss principles and applications of chemistry including forms of matter and energy, properties of substances and reactivity, thermochemistry, periodic properties of elements and compounds, and atomic-molecular structure and bonding. Throughout the course we will identify and solve problems in the physical and biological sciences that are related to these topics. This course is designed for science majors and minors including the following areas: chemistry, biology, physics, environmental science, geology, nutrition, and several pre-professional programs. The course will cover Chapters 1-9 of the textbook, and will emphasize atomic-molecular structure, chemical reactivity and stoichiometry; along with the development of conceptual understanding and problem-solving skills.



Class time and slides are not envisioned to summarize the content of the textbook or substitute independent study of the material. To fully take advantage of class, you will need to **do the assigned readings before class** (see Schedule below) and to participate actively during class. Class time will be used to discuss those aspects of the material that may be more challenging, and to address students' questions and comments. The design of this course is intended to create a collaborative learning environment where participants will interact and work together in small groups (3-4 members) for short periods during each class meeting. Besides being an effective learning strategy, group work in this class counts towards your in-class assignments grade (see Grading below).

### Course Objective

This course will cover forms of matter and energy, chemical and physical properties, thermochemistry, atomic-molecular structure and bonding, and periodic properties of elements. This course intends to help you develop skills to process, analyze and synthesize pertinent chemical information and to draw inferences and make predictions about chemical systems. In addition to increasing your scientific knowledge base and your scientific reasoning skills, this course will help you enhance your ability to explain chemical phenomena based on scientifically sound principles.

### Expected Learning Outcomes

After completing this course, students are expected to:

- Describe and explain the relationship between structure and reactivity.
- Develop interest in science, and in learning science and about science, and they will value scientific literacy for its impact on society.
- Develop critical thinking and problem-solving skills as applied to chemistry and science in general.
- Be able to interpret the chemical significance of new information and ideas acquired in the future.
- Develop and use teamwork abilities (among them communication).
- Be able to identify important resources for their own subsequent learning.

### Instructional technologies

**Blackboard:** This class has its own Blackboard site. You are responsible for frequently checking the site for announcements and other information such as but not limited to homework assignments. I will report grades (quizzes, assignments, exams, etc) through Blackboard. You are responsible to verify the accuracy of the grades entered and for reporting errors to the instructor within one week of the posting date.

**Email:** The best way to reach the instructor outside class is via email. If you do not get a response within 48 hours, you should assume your email was not delivered and you should try to contact the instructor again. If the answer to your inquiry has been posted in Blackboard, you will be simply directed to our page. I welcome your emails; however, given the number of



students in this class I ask you to make sure the answer to your question is not on Blackboard before emailing me.

**Remote voting system:** iClicker remotes must be purchased or reused from a previous class. Students will register their iClicker remote during lecture near the beginning of the semester under the guidance of the professor. iClicker questions are used in every lecture and points are received for both participation and correct answers using the iClicker remote. There is no make-up for clicker questions. We shall start using “clickers” the second week of classes; if you have any doubts please wait until after the first day of class. The iclicker app “Reef polling” will be available as well, so you can use your smart phone, tablet, or laptop during class.

**Homework assignments:** This course uses Blackboard to post homework assignments.

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## GRADING POLICY

Percentages are distributed as shown below. Letter grades are calculated from percentages according to the cutoff values shown below. **There is no extra credit in this course.**

Letter Grade	Final Average
A	90 – 100%
B	80 – 89%
C	70 – 79%
D	60 – 69%
F	< 60%

A weighted percentage will be used to determine the overall course grade.

Homework	9%
In-class Assignments	9%
Exam 1	10%
Exam 2	13.5%
Exam 3	13.5%
Final Exam (comprehensive)	25%
Laboratory	20%

To pass this course with a **grade of A, B or C or better**, I recommend **studying two hours for each hour spent in lecture**. Lecture will take three hours per week, so six hours of studying time (including reading, practice problems, and homework) will be necessary to successfully complete this course.

The laboratory portion of this course counts for 20% of the overall grade. In order to make a C or higher in this course (which is the prerequisite for CHEM 1412), you must earn at least a C in both the lecture portion of the course **AND** in the laboratory portion of the course. See your laboratory syllabus for details regarding the laboratory portion of the course.



## EVALUATION ITEMS

**In-class assignments:** In-class assignments are intended to help you prepare for exams. Diligence in class typically saves time when studying for exams. Students are expected to come to class prepared and to actively contribute to the group discussions and group work. In-class assignments include: **(A) quizzes, (B) group work and (C) clicker questions.** These three items are weighted equally. (A) There will be between five and ten quizzes (may be unannounced), all weighted equally. (B) Students will form groups of three or four members; the composition of the groups is up to the students and may vary from class to class. Groups of less than three members or more than four will not receive credit. Each group will turn in a single worksheet at the end of most class days. Each student must sign their name clearly on this worksheet. I will collect at least eight worksheets over the semester. All worksheets are weighted equally. I will operate under the assumption that learners will not require excessive encouragement to actively engage in the discussions (If you think this may constitute an obstacle please see the instructor to discuss preventive/corrective strategies). (C) We will have clicker questions most class days. Failure to click because of lack of attention, tardiness or not having the device will lead to students not receiving credit for the day/question.

**Homework assignments:** Approach these assignments as a learning tool, not as chore. There will be **four homework assignments** during the semester. Each homework assignment will consist of a combination of multiple choice and short answer items. Deadlines will be posted on Blackboard and during class. Incomplete assignments will receive partial credit (only answered questions will be graded). No homework scores will be dropped. Final homework grade will be the average of all individual homework grades. **Late submissions are not accepted and extensions are not granted.**

**Exams:** Combination of multiple choice and short answer items; 45-50 minutes in length. For these exams you will need to bring: #2 pencils, eraser, **calculator (recommended type is TI-30X, and cannot use smart phone, tablet, or laptop).** Please, arrive 5-10 min before test time. The most important concepts, as stressed primarily in the lectures and in the homework, will predominate in the exam material. You will not be allowed to bring any notes into the examination. Any information brought into the examination through notes, through programmable calculators or cell phones constitutes a violation of the university's policy of academic integrity (see below).



COURSE SCHEDULE (*tentative*)

Week	Day	Date	Topic
1	1	23-Aug	Syllabus and Chapter 1: Matter and Measurement
	2	25-Aug	Chapter 1: Matter and Measurement
2	3	28-Aug	Chapter 1: Matter and Measurement
	4	30-Aug	Chapter 1: Matter and Measurement
	5	1-Sep	Chapter 2: Atoms, Molecules, and Ions
3	6	4-Sep	<b>Labor Day (No Classes)</b>
	7	6-Sep	Chapter 2: Atoms, Molecules, and Ions
	8	8-Sep	Chapter 2: Atoms, Molecules, and Ions
4	9	11-Sep	Chapter 3: Chemical Reactions and Stoichiometry
	10	13-Sep	Chapter 3: Chemical Reactions and Stoichiometry
	11	15-Sep	<b>Exam 1 (Chapters 1, 2)</b>
5	12	18-Sep	Chapter 3: Chemical Reactions and Stoichiometry
	13	20-Sep	Chapter 3: Chemical Reactions and Stoichiometry
	14	22-Sep	Chapter 3: Chemical Reactions and Stoichiometry
6	15	25-Sep	Chapter 4: Reactions in Aqueous Solutions
	16	27-Sep	Chapter 4: Reactions in Aqueous Solutions
	17	29-Sep	Chapter 4: Reactions in Aqueous Solutions
7	18	2-Oct	Chapter 4: Reactions in Aqueous Solutions
	19	4-Oct	Chapter 5: Thermochemistry
	20	6-Oct	<b>Exam 2 (Chapters 3, 4)</b>
8	21	9-Oct	Chapter 5: Thermochemistry
	22	11-Oct	Chapter 5: Thermochemistry
	23	13-Oct	Chapter 5: Thermochemistry
9	24	16-Oct	Chapter 6: Electronic Structure of Atoms
	25	18-Oct	Chapter 6: Electronic Structure of Atoms
	26	20-Oct	Chapter 6: Electronic Structure of Atoms
10	27	23-Oct	Chapter 6: Electronic Structure of Atoms
	28	25-Oct	Chapter 7: Periodic Properties of the Elements
	29	27-Oct	Chapter 7: Periodic Properties of the Elements
11	30	30-Oct	Chapter 7: Periodic Properties of the Elements
	31	1-Nov	Chapter 8: Basic Concepts of Chemical Bonding
	32	3-Nov	Chapter 8: Basic Concepts of Chemical Bonding
12	33	6-Nov	Chapter 8: Basic Concepts of Chemical Bonding
	34	8-Nov	<b>Exam 3 (Chapters 5, 6, 7)</b>
	35	10-Nov	Chapter 8: Basic Concepts of Chemical Bonding
13	36	13-Nov	Chapter 8: Basic Concepts of Chemical Bonding
	37	15-Nov	Chapter 9: Molecular Geometry and Bonding Theories
	38	17-Nov	Chapter 9: Molecular Geometry and Bonding Theories
14	39	20-Nov	Chapter 9: Molecular Geometry and Bonding Theories
	40	22-Nov	<b>Thanksgiving Holidays</b>
	41	24-Nov	<b>Thanksgiving Holidays</b>
15	42	27-Nov	Chapter 9: Molecular Geometry and Bonding Theories
	43	29-Nov	Chapter 9: Molecular Geometry and Bonding Theories
	44	1-Dec	Review
16	45	4-Dec	<b>Final Exam (Chapters 1-9): Monday 6 - 8 p.m.</b>
17	--	11-Dec	Grades Due



## ATTENDANCE POLICY

Attendance to all class meetings is expected; **there is no credit associated with attendance.** Attendance will be taken via clickers. Attendance is not taken for credit but to advise students in case they are not performing at the level of their own expectations. Successful completion of this course will require attendance as you will be tested on material covered in class. Homework sets will also be based on material covered in class. I will try to accommodate extenuating circumstances. If you are absent due to an activity such as a conference, religious holiday, sport event, etc. you must inform me ahead of time. Students are responsible for all material covered during an absence. If missed, exams will only be made up under exceptional circumstances.

This course is designed to develop your critical thinking skills with regards to chemical processes. Expanding your critical thinking skills is a difficult task; the recommended path to success is working through problems. Be an active learner and participate in lectures and office hours. Please make sure that you understand how to work the homework problems for they will likely be on the examinations; ask for help if necessary.

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## UNIVERSITY POLICIES

Please follow the links below to review University Policies that apply to this course:

<http://www.shsu.edu/syllabus/>

<http://www.shsu.edu/dept/dean-of-students/policies/>

### Academic Dishonesty

All students are expected to engage in all academic pursuits in a manner that is above reproach. Students are expected to maintain honesty and integrity in the academic experiences both in and out of the classroom. Any student found guilty of dishonesty in any phase of academic work will be subject to disciplinary action. The University and its official representatives may initiate disciplinary proceedings against a student accused of any form of academic dishonesty including but not limited to, cheating on an examination or other academic work which is to be submitted, plagiarism, collusion and the abuse of resource materials. For a complete listing of the university policy, see: Dean of Student's Office <http://www.shsu.edu/dept/dean-of-students/index.html>.

### Students with Disabilities

It is the policy of Sam Houston State University that individuals otherwise qualified shall not be excluded, solely by reason of their disability, from participation in any academic program of the university. Further, they shall not be denied the benefits of these programs nor shall they be subjected to discrimination. Students with disabilities that might affect their academic performance should register with the Office of Services for Students with Disabilities located in the Lee Drain Annex (telephone 936-294-3512, TDD 936-294-3786, and email [disability@shsu.edu](mailto:disability@shsu.edu)). They should then make arrangements with their individual instructors so that appropriate strategies can be considered and helpful procedures can be developed to ensure that participation and achievement opportunities are not impaired.

SHSU adheres to all applicable federal, state, and local laws, regulations, and guidelines with respect to providing reasonable accommodations for students with disabilities. If you have a



disability that may affect adversely your work in this class, then I encourage you to register with the SHSU Services for Students with Disabilities and to talk with me about how I can best help you. All disclosures of disabilities will be kept strictly confidential. NOTE: No accommodation can be made until you register with the Services for Students with Disabilities. For a complete listing of the university policy, see:

<http://www.shsu.edu/dotAsset/7ff819c3-39f3-491d-b688-db5a330ced92.pdf>

### **Student Absences on Religious Holy Days Policy**

Section 51.911(b) of the Texas Education Code requires that an institution of higher education excuse a student from attending classes or other required activities, including examinations, for the observance of a religious holy day, including travel for that purpose. Section 51.911 (a) (2) defines a religious holy day as: “a holy day observed by a religion whose places of worship are exempt from property taxation under Section 11.20....” A student whose absence is excused under this subsection may not be penalized for that absence and shall be allowed to take an examination or complete an assignment from which the student is excused within a reasonable time after the absence.

University policy 861001 provides the procedures to be followed by the student and instructor. A student desiring to absent himself/herself from a scheduled class in order to observe (a) religious holy day(s) shall present to each instructor involved a written statement concerning the religious holy day(s). The instructor will complete a form notifying the student of a reasonable timeframe in which the missed assignments and/or examinations are to be completed. For a complete listing of the university policy, see:

<http://www.shsu.edu/dept/academic-affairs/documents/aps/students/861001.pdf>

### **Visitors in the Classroom**

Only registered students may attend class. Exceptions can be made on a case-by-case basis by the professor. In all cases, visitors must not present a disruption to the class by their attendance. Students wishing to audit a class must apply to do so through the Registrar's Office.

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### **Notes**

- 1) The course syllabus is intended to provide students with basic information concerning the course. The syllabus can be viewed as a “blueprint” for the course; changes in the syllabus can be made and students will be informed of any substantial changes concerning exams, grading or attendance policy and/or changes to reading or homework assignments.